

Journal 30: Naming Practice



Cs_2O	Cesium oxide
Na_3N	Sodium nitride Na^{1+} N^{3-}
$\text{Be}(\text{C}_2\text{H}_3\text{O}_2)_2$	Beryllium acetate Be^{2+} $\text{C}_2\text{H}_3\text{O}^-$
$\text{Zn}_3(\text{PO}_4)_2$	Zinc phosphate
$(\text{NH}_4)_4\text{C}$	Ammonium Carbide NH_4^+ C^{-4}

Molecular Compounds

Naming Molecular Compounds & Writing Names from Formulas

1. Both elements are nonmetals.
2. Use prefixes to indicate how many atoms are present
3. Naming the first element: prefix with its full name;
Do not use “mono” when naming the first element
4. Name the second element: always use a prefix, then element name, and change the ending to -ide

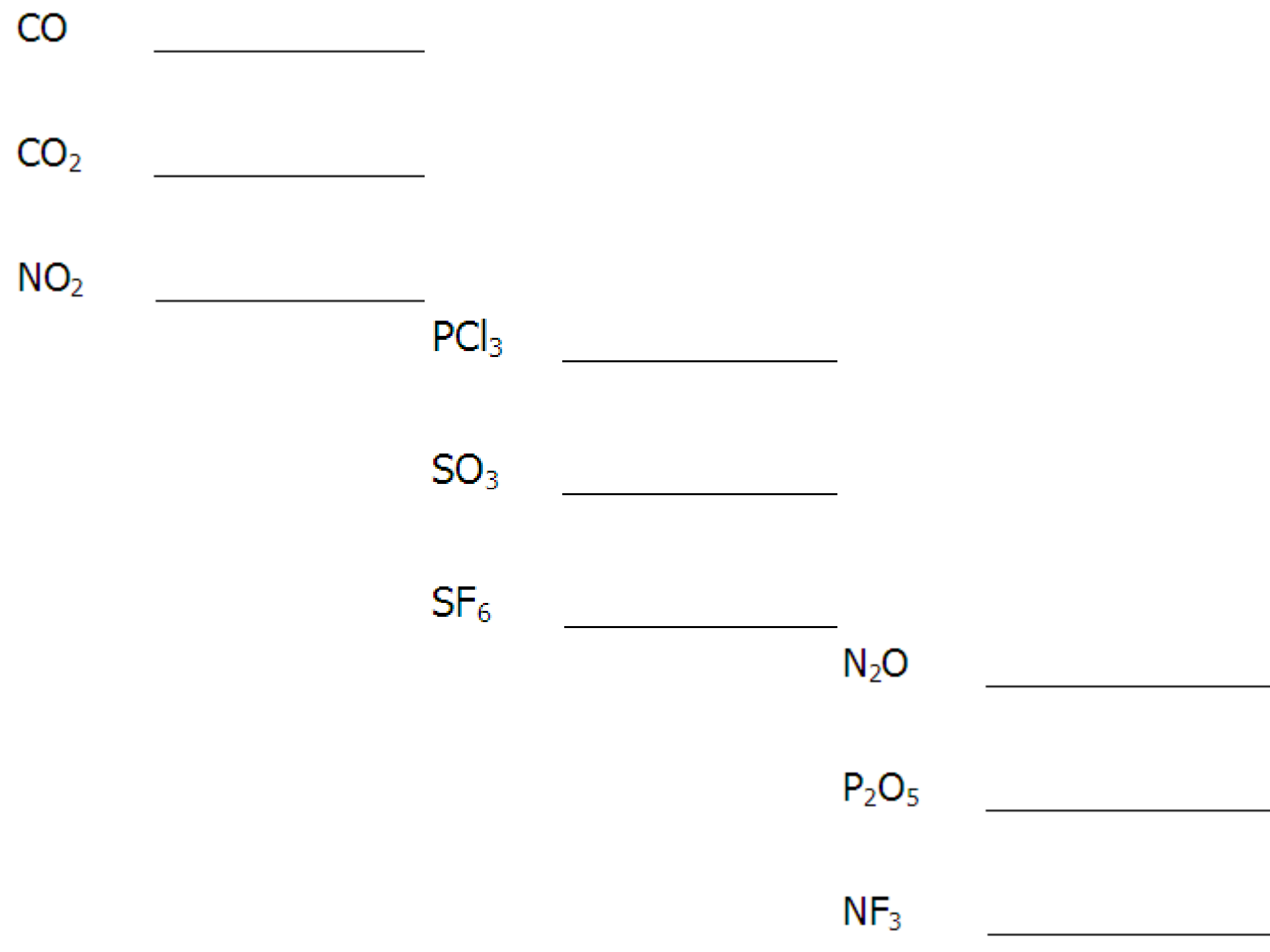
Example: CS_2

Elements: carbon & 2 sulfur

Name: carbon *disulfide*

Prefixes:

1: mono	2: di	3: tri	4: tetra	5: penta
6: hexa	7: hepta	8: octa	9: nano	10: deca



Writing Formulas from Names

1. Identify the elements
2. Identify prefixes for each element
3. Charges do not matter for molecular formulas!!

diphosphorous monosulfide

sulfur tetrafluoride

nitrogen monoxide

carbon monoxide

nitrogen triiodide

phosphorous hexabromide

diphosphorous trioxide

carbon tetrabromide

dichlorine heptoxide

Molecular Worksheet

- Molecular Worksheet
- Homework - Pg. 231 #4a-e; pg 235 #2; pg 251 #10, 11