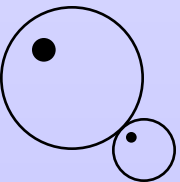
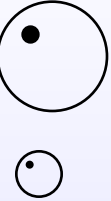
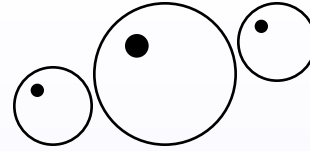
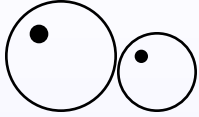
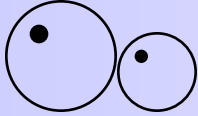


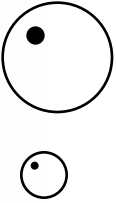
Reaction Classification





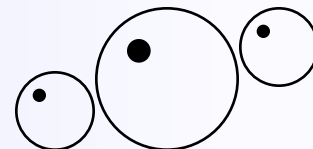
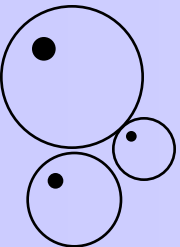
Classifying Reactions

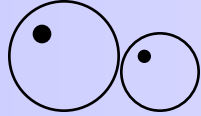
To classify a reaction we need to look at the reaction equation.



We are looking for:

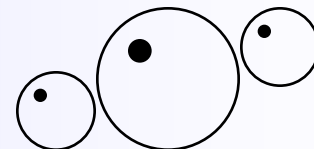
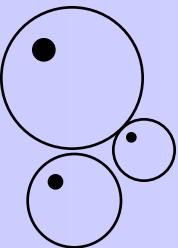
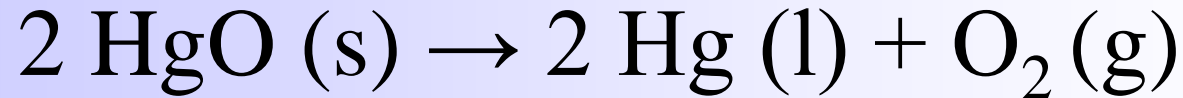
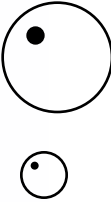
- specific reactants and products
- states
- relationship between reactants and products





Decomposition Reactions

Compound is broken into simpler pieces

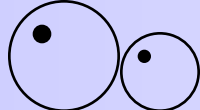




Synthesis Reactions

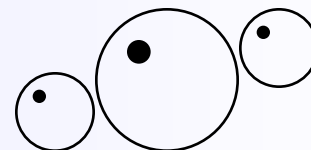
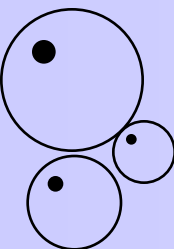
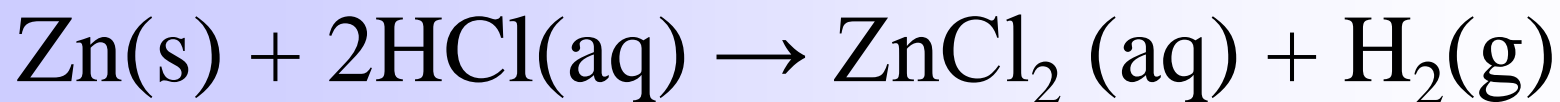
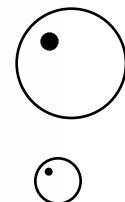
Two or more reactants combine to form one product

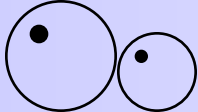




Single Displacement Reactions

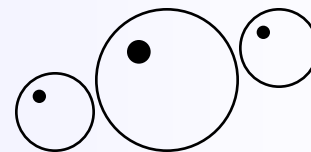
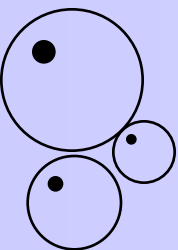
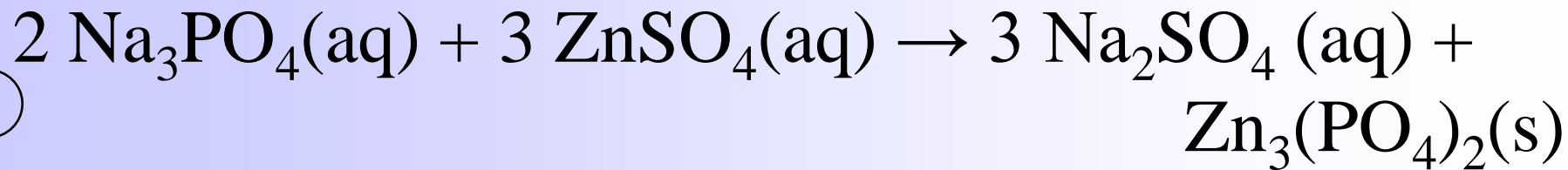
One element in a compound is kicked out and replaced by another





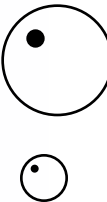
Double Displacement Reactions

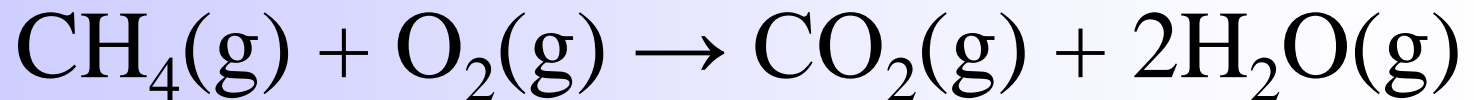
- Ion swap
- May form precipitate (*use solubility table*)
- May form water

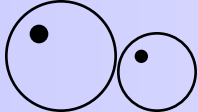




Combustion Reactions

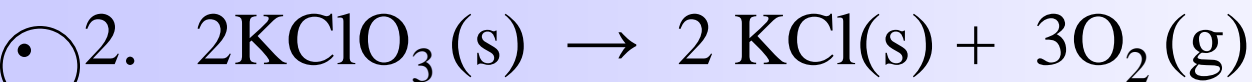
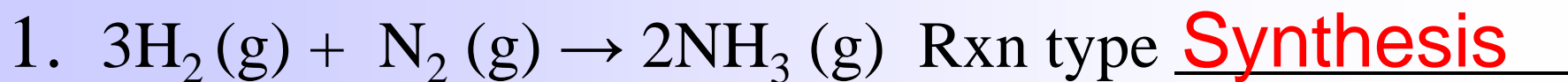
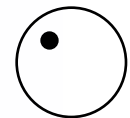
- Oxygen must be a reactant
 - Heat and flame are produced
- 



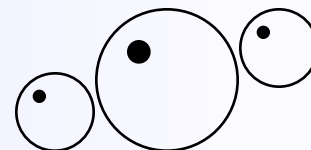
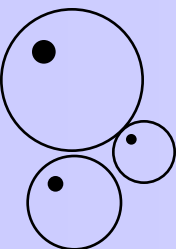


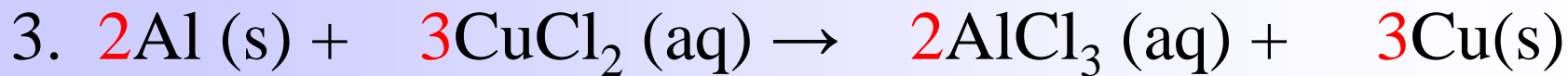
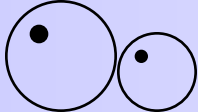
Lets practice... Reaction Types

Decomposition; Synthesis; Single Displacement
Double Displacement; Combustion



Rxn type Decomposition

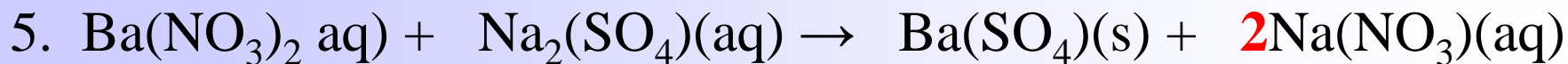




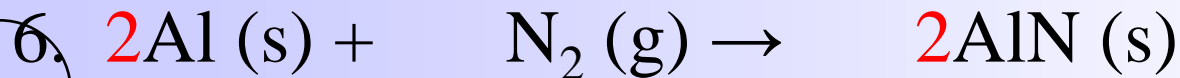
Rxn type Single Rep.



Rxn type Decomposition



Rxn type Double Rep.



Rxn type Synthesis

